Common elements:

H - hydrogenNi - nickelLi - lithiumCu - copperNa - sodiumAg - silverK - potassiumAu - goldMg - magnesiumAl - aluminium

Al – aluminium Ca - calcium Zn - zinc Ti- titanium Hg - mercury C – carbon Si – silicon Cr - chromium S – sulphur Mo - molybdenum Cl – chlorine W - wolfram, tungsten Pb - lead Mn - manganese Ne - neon Cd - cadmium Xe - xenon

N- nitrogen Zn- zinc O- oxygen Sn- tin

Fe - iron Ph — phosphorus

Co - cobalt I – iodine

Oxides

Czech terminology - suffixes: - ný, - natý, - itý, -ičitý, -ečný/ičný, ový, -istý, -ičelý.

- 1. -ný např. Li2O, oxid lithný
- 2. -natý např. MgO, oxid hořečnatý
- 3. -itý např. Al2O3, oxid hlinitý
- 4. -ičitý např. CO2, oxid uhličitý
- 5. -ičný/-ečný např. P2O5, oxid fosforečný
- 6. -ový např. MoO3, oxid molybdenový
- 7. -istý např. Cl2O7, oxid chloristý
- 8. -ičelý např. OsO4, oxid osmičelý

English terminology - prefixes: mono-, di-, tri-, tetra-, penta-, hexa-, hepta-, octa-.

Examples of oxides

The following table gives examples of commonly encountered oxides. Only a few representatives are given, as the number of polyatomic ions encountered in practice is very large.

Name	Formula	Found/Usage
Water (hydrogen oxide)	H₂O	Common solvent, required by carbon-based life
Nitrous oxide	N ₂ O	Laughing gas, anesthetic (used in a combination with diatomic oxygen (O_2) to make nitrous oxide and oxygenanesthesia), produced by nitrogen-fixing bacteria, nitrous, oxidizing agent in rocketry, aerosol propellant, recreational drug, greenhouse gas. Other nitrogen oxides such as NO ₂ (nitrogen dioxide), NO (nitrogen oxide), N ₂ O ₃ (dinitrogen trioxide) and N ₂ O ₄ (dinitrogen tetroxide) exist, particularly in areas with notable air pollution. They are also strong oxidisers, can add nitric acid to acid rain, and are harmful to health.
Silicon dioxide	SiO ₂	Sand, quartz
<u>Iron(II,III)</u> <u>oxide</u>	Fe ₃ O ₄	Iron ore, rust, along with iron(III) oxide (Fe ₂ O ₃)
Aluminium oxide	Al ₂ O ₃	Aluminium ore, alumina, corundum, ruby (corundum with impurities of chromium).

Zinc oxide	ZnO	Required for <u>vulcanization</u> of <u>rubber</u> , additive to <u>concrete</u> , <u>sunscreen</u> , <u>skin</u> <u>care</u> lotions, <u>antibacterial</u> and <u>antifungal</u> properties, <u>food additive</u> , white <u>pigment</u> .
<u>Carbon</u> <u>dioxide</u>	CO ₂	Constituent of the <u>atmosphere of Earth</u> , the most abundant and important <u>greenhouse gas</u> , used by <u>plants</u> in <u>photosynthesis</u> to make <u>sugars</u> , product of biological processes such as <u>respiration</u> and chemical reactions such as <u>combustion</u> and <u>chemical</u> <u>decomposition</u> of <u>carbonates</u> . CO or <u>Carbon monoxide</u> exists as a product of <u>incomplete combustion</u> and is a highly <u>toxic gas</u> .
<u>Calcium</u> <u>oxide</u>	CaO	Quicklime (used in construction to make mortar and concrete), used in self-heating cans due to exothermic reaction with water to produce calcium hydroxide, possible ingredient in Greek fire and produces limelightwhen heated over 2,400 °Celsius.